

Practice and Revision - Carbon and Its Compounds

Chapter 17: Carbon and Its Compounds · Topic 2

Key Concept

Carbon exists in different allotropes (different structural forms), like diamond (hard, transparent) and graphite (soft, used in pencils).

Carbon's ability to form four bonds allows it to create complex chains and rings, forming the basis of organic chemistry.

Carbon compounds are essential for life, forming molecules like carbohydrates, proteins, and fats found in living organisms.

Understanding carbon's unique bonding properties helps explain the vast diversity of organic compounds essential to life and industry.

Quick Revision Notes

- Carbon has different allotropes.
- Diamond is hard and transparent.
- Graphite is soft, used in pencils.
- Carbon forms four bonds.
- This allows complex chains and rings.
- This is the basis of organic chemistry.
- Carbon compounds form life's essential molecules.
- Examples: carbohydrates, proteins, fats.

Practice Questions

- 1 Research diamond and graphite as carbon allotropes.
- 2 Discuss why carbon can form so many compounds.
- 3 Research an organic compound, like glucose.
- 4 Explain why carbon is essential for life.
- 5 Discuss the properties of diamond versus graphite.
- 6 Research another use of a carbon-based material.