

Neutralisation Reaction

Chapter 2: Exploring Substances: Acidic, Basic, and Neutral · Topic 2

Key Concept

Farmers use neutralisation when they add lime (a base) to acidic soil to make it more suitable for growing crops. In factories, neutralisation is used to safely treat acidic or basic waste before it is released into the environment, preventing damage to rivers and ecosystems. The general pattern of this reaction is: $\text{Acid} + \text{Base} \rightarrow \text{Salt} + \text{Water}$, and the type of salt formed depends on which specific acid and base were used.

Quick Revision Notes

- Neutralisation: $\text{Acid} + \text{Base} \rightarrow \text{Salt} + \text{Water}$.
- Antacids are mild bases used to neutralise excess stomach acid.
- Farmers add lime (a base) to neutralise acidic soil.
- Industries use neutralisation to safely treat acidic/basic waste.

Practice Questions

- 1 Explain why antacids are used to treat acidity.
- 2 Why might a farmer add lime to acidic farmland?
- 3 Write the general word equation for a neutralisation reaction.